INTRODUCTION TO TITRATION

Name: ———	— Date: ———	Station #:
	ractice dilution, and the use of titration to	For volumetric analysis and to familiarize with all the re.
Procedure: As	in CHEM 1105 lab manual, pages	<u>.</u>
Observations:		
Part I:		
Part II:		
1 411 11.		
Data:		
Part I:		
Concentration of	pipetting solution =	
Volume of solution	on pipetted =	
Total volume of d	ilute solution prepared =	

Part II:	
Molarity of NaOH:	
Volume of HCl pipetted:	

	Run 1	Run 2	Run 3	Run 4
Initial burette vol. (mL)				
Final burette vol. (mL)				
Vol. of NaOH used (mL)				
End Point colour				

Calculations:

Part I:

Calculate the final concentration of the diluted solution

Part II:

1. Calculate the % difference between runs 1 & 2, runs 2 & 3 and runs 1 & 3.

$$\% \ difference = \left| \frac{V_1 - V_2}{\left(\frac{V_1 + V_2}{2}\right)} \right| \times 100\%$$

2. Calculate the concentration of the unknown HCl solution.

Conclusion:

Questions:

1.	Why did you not rinse out your Erlenmeyer flasks with acid before you pipetted the acid in?
2.	Would it matter if your Erlenmeyer flasks were wet (with distilled water) when you pipetted the acid into them?
3.	Would your titration volumes have changed if your beakers had not been dry and before you put the acid or base into them? Why?