

CHEM 1110: DETERMINATION OF ASCORBIC ACID IN VITAMIN C

Date: _____ Name: _____ Lab Day/Time: _____

Objective

To determine the percentage of ascorbic acid in an unknown solution of ascorbic acid.

Procedure

As in Chem 1110 lab manual, pages _____

Observations

Data

Table 1. Mass of KIO_3

Mass of KIO_3 and weighing boat (g)	
Mass of <i>emptied</i> boat (g)	
Mass of KIO_3 emptied into flask (g)	

5. Moles of ascorbic acid present

6. Mass of Ascorbic acid present

7. Mass of unknown solution that was used

8. % Ascorbic acid in unknown solution

9. Calculate the Average

Table 3. Summary of Calculations

Run Number			
Moles of IO_3^- used			
Moles of I_3^- produced			
Moles of Ascorbic Acid present			
Mass of Ascorbic Acid Present			
Mass of Unknown Solution			
% Ascorbic Acid in unknown solution			

Results

Table 4. Summary of Results

Average % Ascorbic Acid in Unknown # _____	
Circle which runs you used to calculate the average?	1 2 3 4

Discussion

Give one possible source of error beyond your reasonable control in the experiment, and explain how and why this would affect your result.

Conclusion

Questions

Attach any questions assigned by your lab instructor.