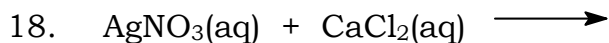
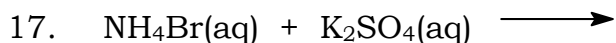
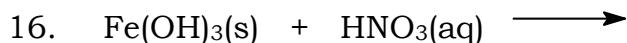
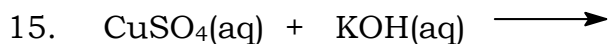
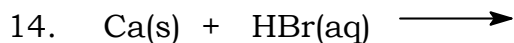
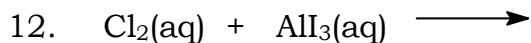
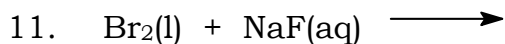
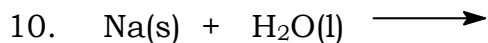
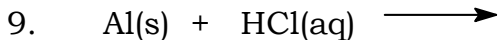
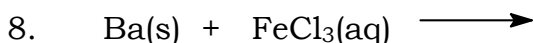
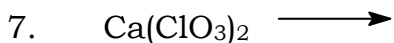
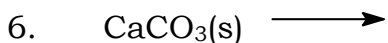
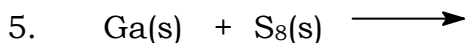
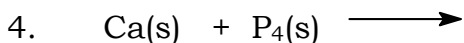
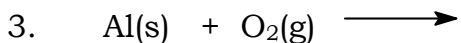
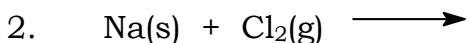
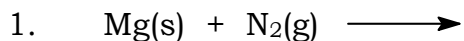


## BALANCING EQUATIONS

A. Complete and balance the following reactions. Give states of the products.



19.  $\text{HCl(aq)} + \text{KHCO}_3\text{(aq)} \longrightarrow$
20.  $\text{HCl(aq)} + \text{Sr(OH)}_2\text{(aq)} \longrightarrow$
21.  $\text{BaCl}_2\text{(aq)} + \text{Na}_2\text{SO}_4\text{(aq)} \longrightarrow$
22.  $\text{Na}_2\text{S(aq)} + \text{Ni(NO}_3)_2\text{(aq)} \longrightarrow$
23.  $\text{C}_2\text{H}_6\text{(g)} + \text{O}_2\text{(g)} \longrightarrow$
24.  $\text{C}_8\text{H}_{18}\text{(l)} + \text{O}_2\text{(g)} \longrightarrow$
25.  $\text{FeCl}_2\text{(aq)} + \text{K}_3\text{PO}_4\text{(aq)} \longrightarrow$

B. Give the net-ionic equations for the reactions 9,10,15,16, and 18 to 22 from part A. The equation should properly be balanced as for as charges and atoms are concerned.