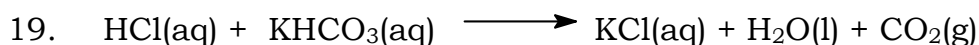
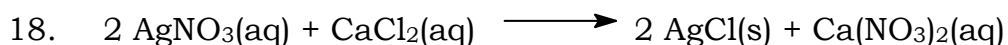
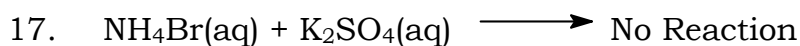
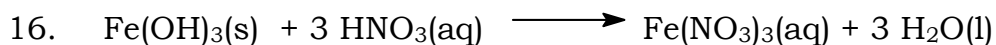
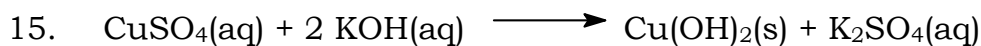
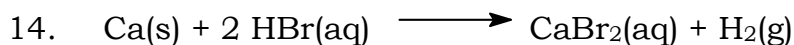
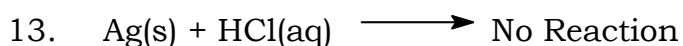
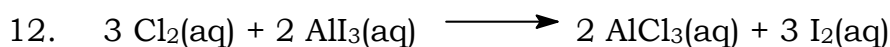
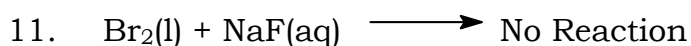
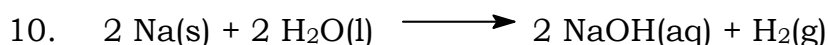
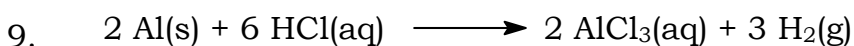
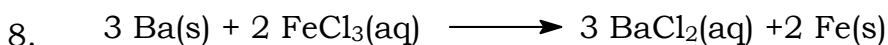
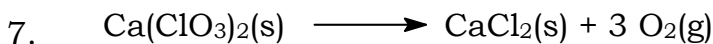
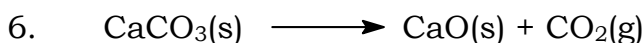
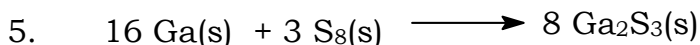
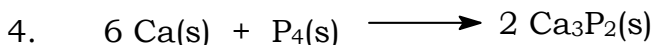
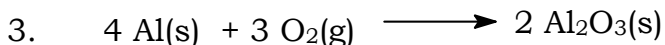
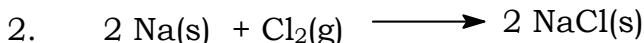
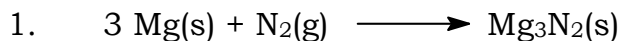


BALANCING EQUATIONS

A. Complete and balance the following reactions. Give states of the products.



20. $2 \text{HCl}(\text{aq}) + \text{Sr}(\text{OH})_2(\text{aq}) \longrightarrow \text{SrCl}_2(\text{aq}) + 2 \text{H}_2\text{O}(\text{l})$
21. $\text{BaCl}_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \longrightarrow \text{BaSO}_4(\text{s}) + 2 \text{NaCl}(\text{aq})$
22. $\text{Na}_2\text{S}(\text{aq}) + \text{Ni}(\text{NO}_3)_2(\text{aq}) \longrightarrow \text{NiS}(\text{s}) + 2 \text{NaNO}_3(\text{aq})$
23. $2 \text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \longrightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$
24. $2 \text{C}_8\text{H}_{18}(\text{l}) + 25 \text{O}_2(\text{g}) \longrightarrow 16 \text{CO}_2(\text{g}) + 18 \text{H}_2\text{O}(\text{l})$
25. $3 \text{FeCl}_2(\text{aq}) + 2 \text{K}_3\text{PO}_4(\text{aq}) \longrightarrow \text{Fe}_3(\text{PO}_4)_2(\text{s}) + 6 \text{KCl}(\text{aq})$

B. Give the net-ionic equations for the reactions 9,10,15,16, and 18 to 22 from part A. The equation should properly be balanced as for charges and atoms are concerned.

