ISOTOPES AND ATOMIC MASS

1. Give the symbol for an isotope that has an atomic number of 19 and a mass number of 40.

2. An isotope of lithium has four neutrons in its nucleus. Write its isotopic notation.

3. By using a mass spectrometer we could find that an atom of the isotope Br-79 has a mass 6.577 times as large as an atom of C-12. What is the atomic mass of the Br-79 isotope?

4. An atom of O-16 is 1.3329096 times more massive than C-12. What is the mass of O-16?

5. Calculate the atomic mass from the following percent abundances and the masses of the isotopes.

	ISOTOPE	**% ABUNDANCE**	**MASS(amu)**
a)	C-12	98.89	
	C-13	1.11	13.0034
b)	Mg-24	78.99	23.9850
	Mg-25	10.00	24.9858
	Mg-26	11.01	25.9826
c)	Cu-63	69.09	62.93
	Cu-65	30.91	64.93
d)	Ga-69	60.27	68.9257
	Ga-71	39.73	70.9249
e)	O-16	99.7587	15.994915
	O-17	0.0374	0.00636
	O-18	0.2039	0.03670
f)	B-10	19.78	10.0129
	B-11	80.22	11.0093